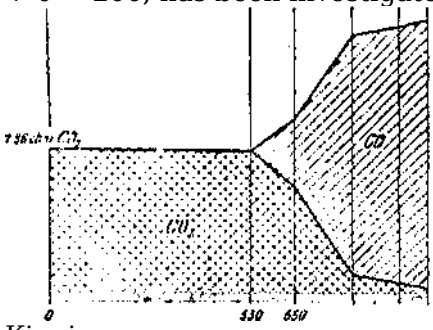
Table B has also boon graphically set out, and one sees from the diagram (Fig. <)) that in gasilieation of carbon by means of steam, at a temperature of about 8(50°, the formation of COg predominates, whilst from that point with increasing tem-poraturo iho formation of carbon monoxide strongly gains the upper hand; and that at 1,125° only the latter gas is formed, and thereat the steam is completely decomposed. It should bo vory specially observed how much the quality of tho gas is damaged by tho passing through of undecomposed stoam, if tho tcmiperaturos in the reaction zone be too low.

Tim action of carbonic at*, id on carbon, which takes place according to tho equation 00.2 + 0 = 200, has been investigated by Boubouard (FiBcher's " Techn. d. Brennst," p. 200)



Kiu. in,

$00 92H WOO0

iwt'H at. o*nncl 7(50 millime.trcB barometer.

at dilYonmt tomporaturoB ; \vhorofrom tho following conditions ol equilibrium havoboon obtained:-—

At ii tnntjMM'ittui'M of Tho ^an mixturo contains

|J«">0'' • * .01 JHT rent. (1< Kj ami US) per (HMI^. CO,

HIM)-' , „ .7 n (XK, „ 1)1 „ C0.»

S»2-l;< . 4 <'<K(„ 00 ,,

CO.

Those roKultiH arc not out in tho diagram (l^ig- 10).

Ono, port'cuvcm front all ihroo cason connidored that the gain in KombuHtiblo ganon IB tho groator, the higher toiuporaturos omployotl. In tho formation of air-gas and the reduction of carbonic acid by means of incandescent carbon, the maximum of formation of carbon monoxide is reached at a temperature of about 1,000\ whilst in the decomposition of water, only at

tlu nut tally (Translator).